

I can perform calculations using scientific notation, significant figures, and unit conversions.

4.1. I can define metric units and prefixes and explain how metric units are used in measurement.

4.2. I can express numbers in scientific notation and standard notation.

4.3. I can perform calculations using significant figures.

4.4. I can convert between units, both metric and US standard.

4.5. I can distinguish between accuracy and precision in a given set of data

Hmwk: DA level 1

$$\textcircled{1} \frac{14 \text{ mm} \times 1 \text{ m}}{1000 \text{ mm}} = \boxed{0.014 \text{ m}}$$

$$\textcircled{7} \frac{45 \text{ ml} \times 11.3 \text{ g}}{1 \text{ ml}} = \boxed{510 \text{ ml}} \\ (508.5)$$

$$\textcircled{2} \frac{35 \text{ kg} \times 1000 \text{ g}}{1 \text{ kg}} = \boxed{3500 \text{ g}}$$

$$\textcircled{8} \frac{100.0 \text{ ml} \times 2.16 \text{ g}}{1 \text{ ml}} = \boxed{216 \text{ g}}$$

$$\textcircled{3} \frac{57 \text{ ml} \times 1 \text{ L}}{1000 \text{ ml}} = \boxed{0.057 \text{ L}}$$

$$\textcircled{9} \frac{5.5 \text{ g} \times 15 \text{ km}}{1 \text{ g}} = \boxed{83 \text{ km}}$$

$$\textcircled{4} \frac{88 \text{ m} \times 100 \text{ cm}}{1 \text{ s} \times 1 \text{ m}} = \boxed{\frac{8800 \text{ cm}}{\text{s}}}$$

$$\textcircled{10} \frac{0.525 \text{ mol} \times 6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol}} = \boxed{3.16 \times 10^{23} \text{ atoms}}$$

$$\textcircled{5} \frac{9.45 \text{ g} \times 1 \text{ L}}{1 \text{ L} \times 100 \text{ ml}} = \boxed{0.00945 \text{ g/ml}}$$

$$\textcircled{11} \frac{2.75 \text{ L} \times 61.2 \text{ g}}{1 \text{ L}} = \boxed{168 \text{ g}}$$

$$\textcircled{6} \frac{3.55 \text{ ml} \times 13.6 \text{ g}}{1 \text{ ml}} = \boxed{48.3 \text{ g}}$$

$$\textcircled{12} \frac{50.0 \text{ ml} \times 0.000245 \text{ g}}{1 \text{ ml}} = \boxed{0.0123 \text{ g}}$$

10/2 Dimensional Analysis Continued

Steps for solving:

- Identify given and needed. (look for questioning words)
- Determine if there are conversion factors given. (equalities)
- Determine if there are one or more units to convert.

given: g/ml need: L

one unit to change

given: g/ml need: kg/L

two units to change
- map out the conversions

given: g/ml need: kg/L

given: km/hr need: cm/sec

(m, min)
- Write down the complete conversion problem. (← POINTS!)
- Solve the problem. Check sigfigs + units.

Showing Labels only - complete the following:

How fast in $\frac{cm}{sec}$ is a kangaroo jumping if its speed is $\frac{68.7 km}{hr}$?

given: $\frac{68.7 km}{1 hr}$ need: $\frac{? cm}{sec}$

(m, min, sec)

$$\frac{68.7 km}{1 hr} \times \frac{1000 m}{1 km} \times \frac{100 cm}{1 m} \times \frac{1 hr}{60 min} \times \frac{1 min}{60 sec} = \frac{1910 \frac{cm}{sec}}{1908.3}$$

$$(68.7 \times 1000 \times 100) \div (60 \times 60) =$$

A snail moves $6.32 \frac{cm}{min}$, what is this in $\frac{km}{hr}$?

$$\frac{6.32 cm}{1 min} \times \frac{1 m}{100 cm} \times \frac{1 km}{1000 m} \times \frac{60 min}{1 hr} = 0.003792 \frac{km}{hr}$$

How many seconds in 28 days? (no sigfig here)

$$\frac{28 days}{1 day} \times \frac{24 hrs}{1 day} \times \frac{60 min}{1 hrs} \times \frac{60 sec}{1 min} = 2419200 sec$$

The density of lead is $13.92 \frac{kg}{L}$, what is that in $\frac{g}{cm^3}$?

* hint: $1 ml = 1 cm^3$