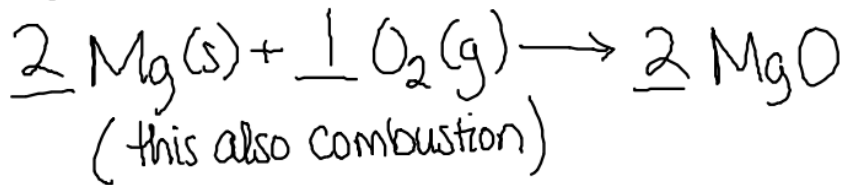


Predicting Products: Chem Rxn Lab.

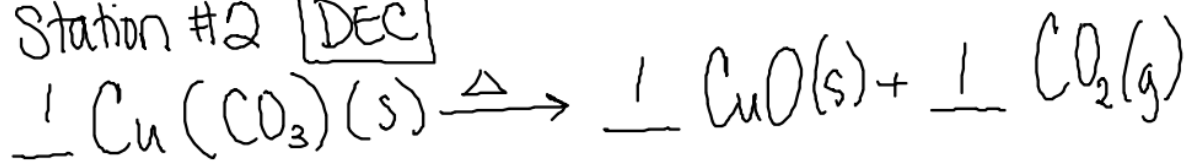
Station #1 DC or Synthesis 2+ 2-



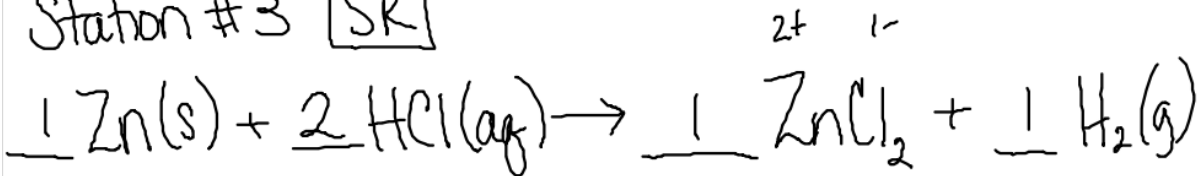
Endothermic takes in heat.

Exothermic gives off heat.

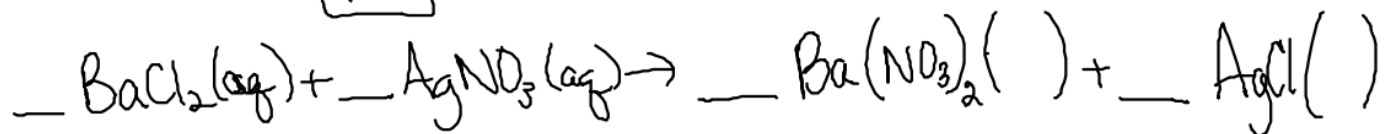
Station #2 DEC



Station #3 SR

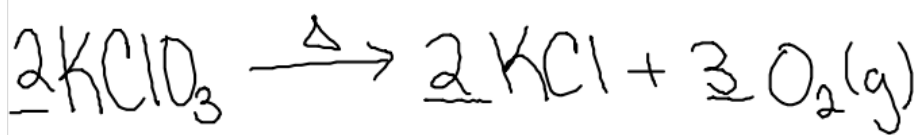
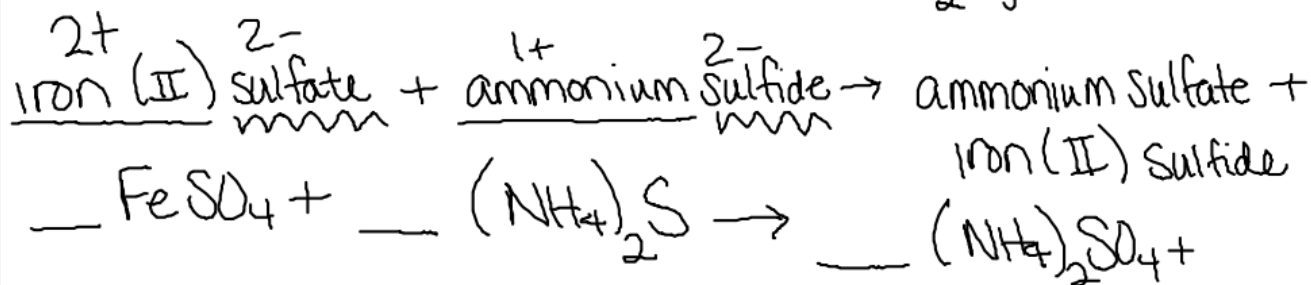
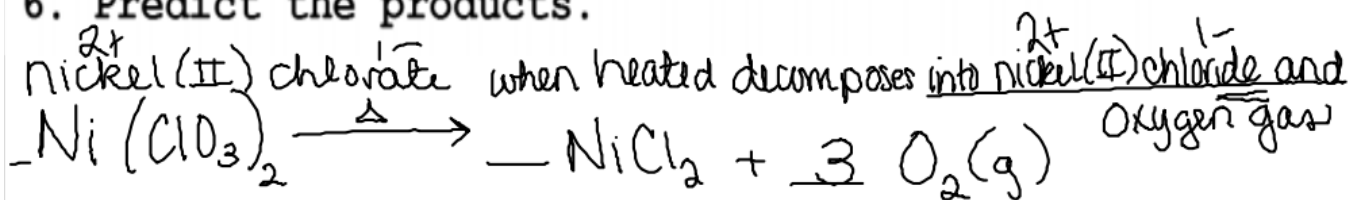


Station #4 DR



The student will be able to:

1. Identify the parts of an chemical reaction.
2. Balance an equation using coefficients.
3. Calculate quantities of a substance in a chemical reaction.
4. Write a chemical equation starting with the chemical names -> written formulas -> balanced reaction
5. Identify the type of reaction.
6. Predict the products.



Reactant: white crystals

Reaction → clear liq. bubbling

Product: clr. liq.

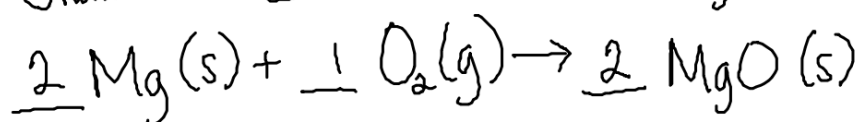
Endothermic — takes in heat

Exothermic — releases heat

Chm. Rxn. Lab

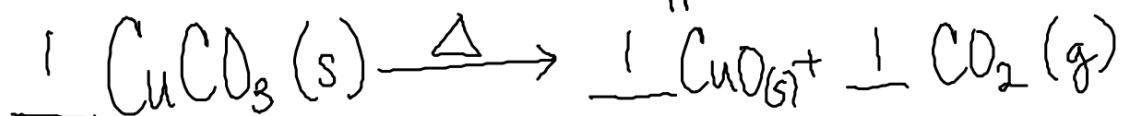
Station #1 [DC]

magnesium²⁺ oxide²⁻



Station #2 [DEC]

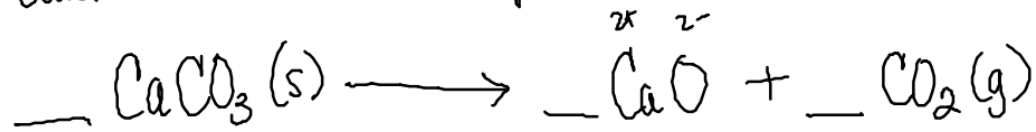
copper(II) oxide + carbon dioxide



Station #4 [DR]

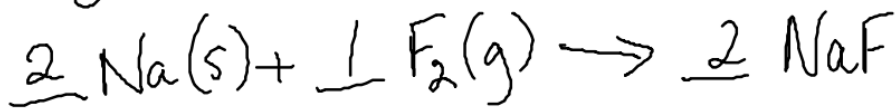


Calcium carbonate decomposes into calcium and carbon dioxide.



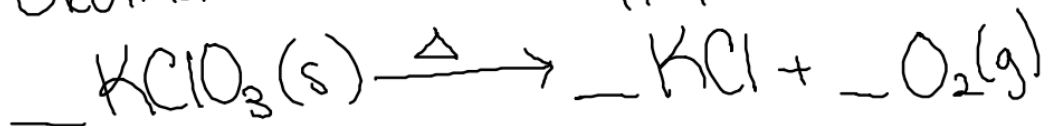
magnesium hydroxide \rightarrow magnesium oxide + water

Sodium reacts with fluorine to form sodium fluoride



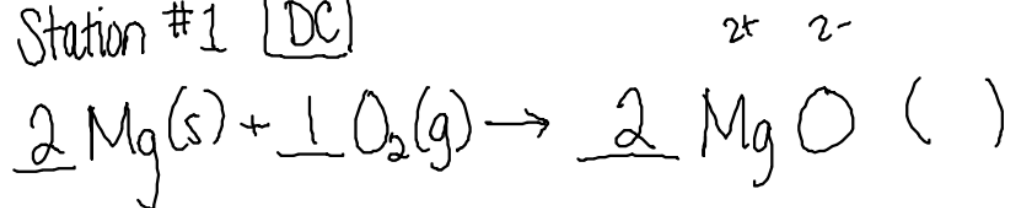
Endothermic - takes in heat

Exothermic - releases heat

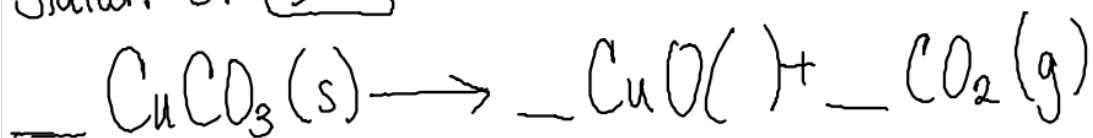


Chem Rxn Lab

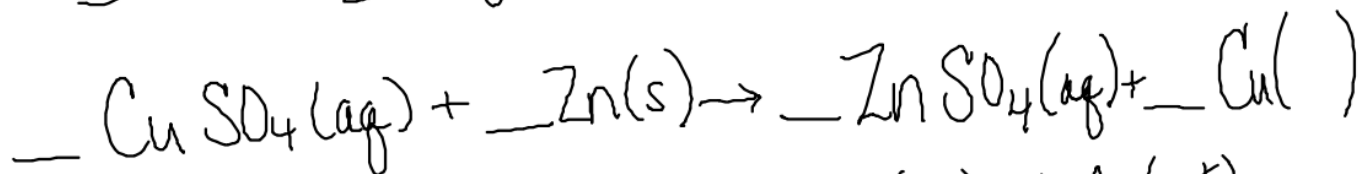
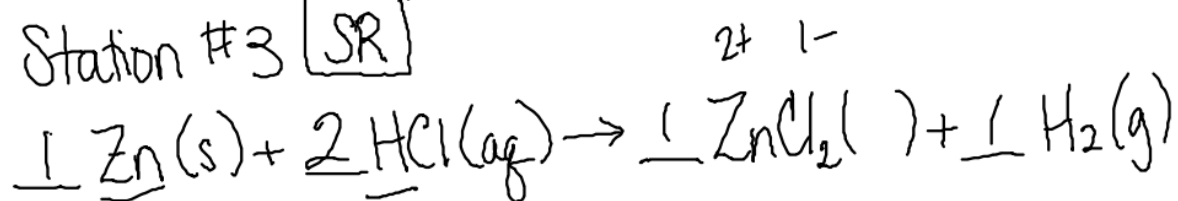
Station #1 [DC]



Station #2 [DEC]



Station #3 [SR]



Station #4 [DR]

