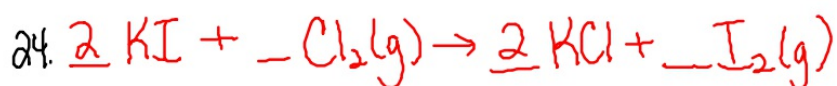
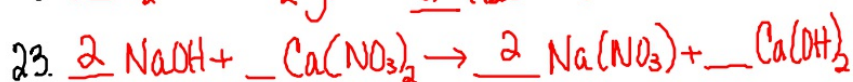
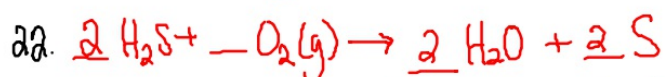
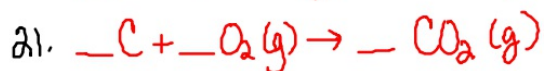


Chem Rxn Review

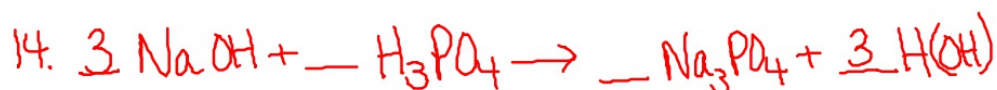
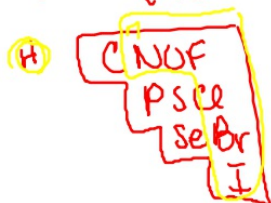
- B 1. 11. 4, 3, 2
C 2. 12. 2, 3, 1, 6
C 3. 13. 3, 2, 6, 1
A 4. 14. 3, 1, 1, 3
B 5. 15. 4, 5, 2
A 6. 16. 1, 4, 1, 4
B 7. 17. 2, 3, 1, 3
B 8. 18. 2, 1, 2
A 9. 19. 3, 2, 1, 3
C 10.



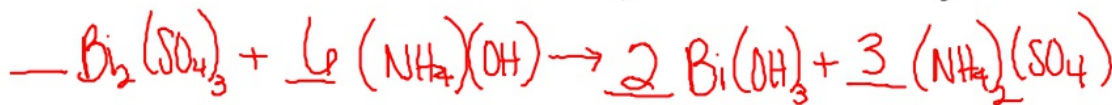
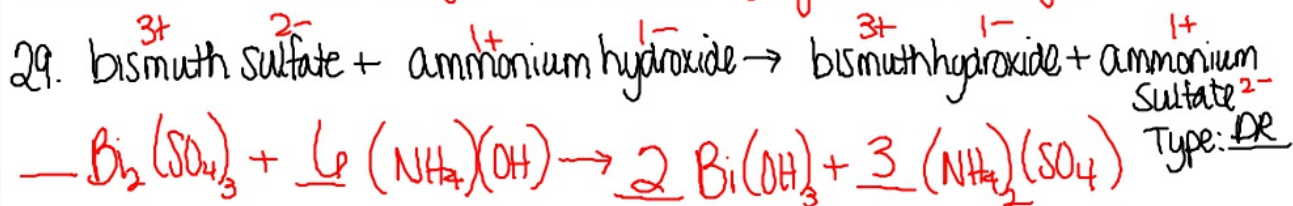
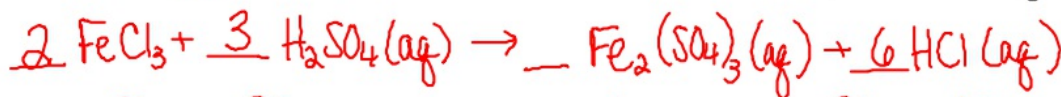
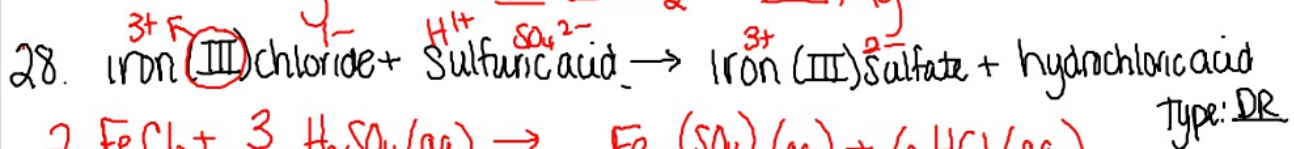
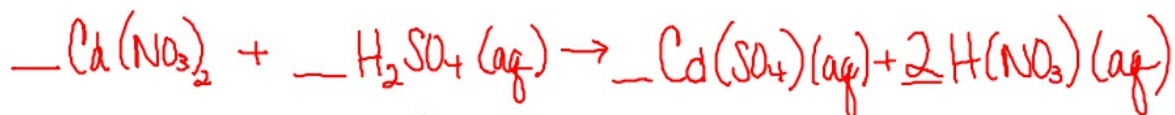
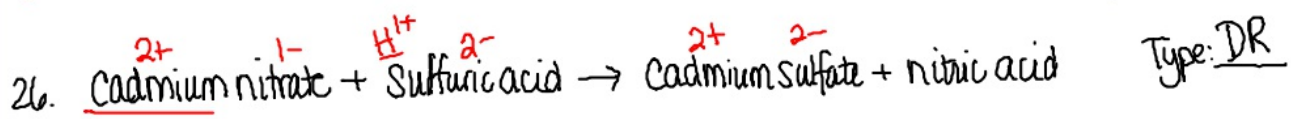
-ide (No oxygen)

(aq) aqueous

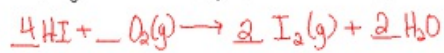
H+_(aq) is an acid



DC	direct combination	Single + single	
SR	single replacement	Single + compound	(like replaces like) + for + - for -
DR	double replacement	compound + compound	(+ switch)
DEC	decomposition	compound	Know the rules — binary chlorates carbonates hydroxides
COMB	combustion	+ O ₂ (g)	



30. hydrogen iodide + oxygen gas \rightarrow iodine gas + water Type: SR



31. nitrogen gas + hydrogen gas \rightarrow ammonia Type: DC



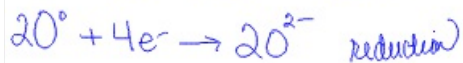
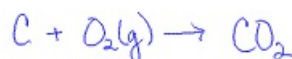
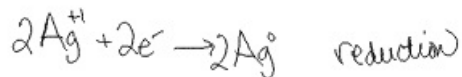
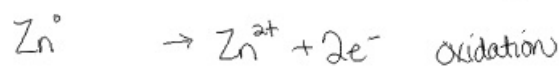
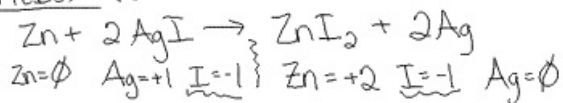
32. aluminum hydroxide (heated) $\xrightarrow{\Delta}$ aluminum oxide + water Type: DEC



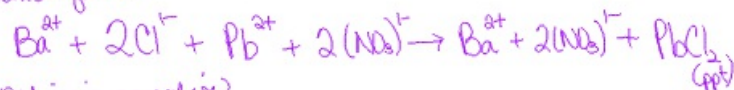
33. burning methane \rightarrow carbon dioxide + water Type: Comb



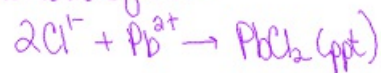
REDOX (DC or SR)



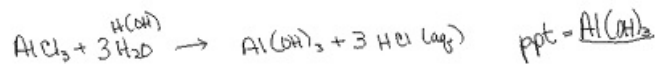
ionic equation



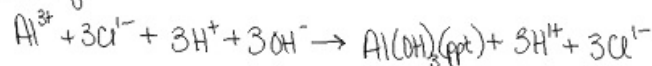
net ionic equation



aluminum chloride + water \rightarrow ~~aluminum hydroxide~~ ~~hydrochloric acid~~



ionic eq:



net ionic:

