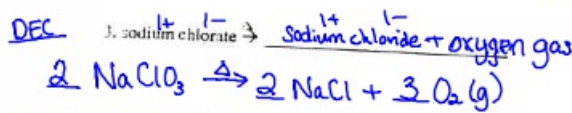
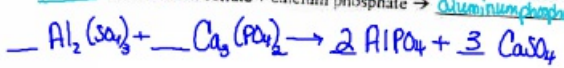


Name: _____ Period: _____ Date: _____

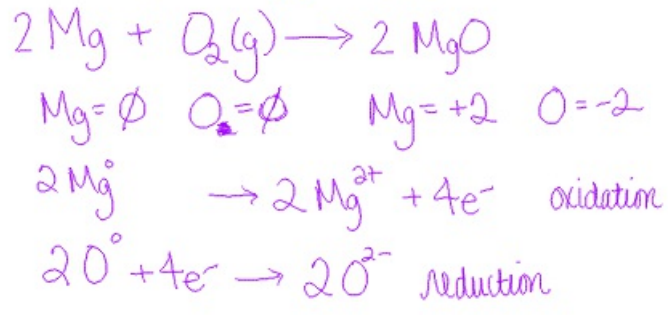
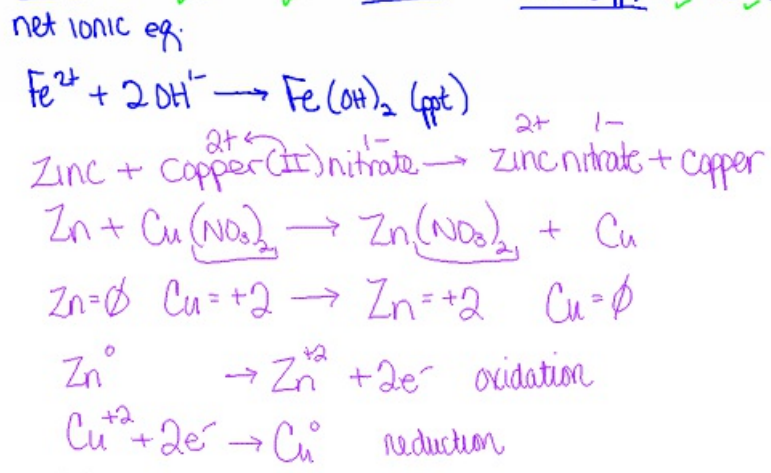
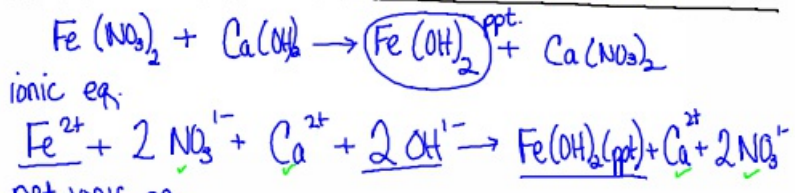
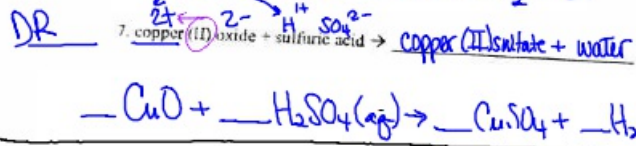
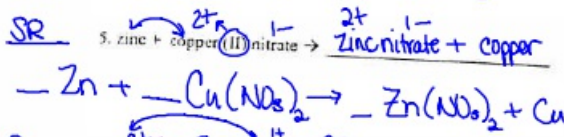
Homework: Types of Reactions

- Identify the type of reactions.
- $1 \text{ CuBr}_2 + 2 \text{ NaI} \rightarrow \text{CuI}_2 + 2 \text{ NaBr}$
 - $2 \text{ KI} + \text{Br}_2(\text{g}) \rightarrow 2 \text{ KBr} + \text{I}_2(\text{g})$
 - $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow 6 \text{ C} + 6 \text{ H}_2\text{O}$
 - $2 \text{ NaF} \rightarrow 2 \text{ Na} + \text{F}_2(\text{g})$
 - $\text{Si} + \text{O}_2(\text{g}) \rightarrow \text{SiO}_2$
 - $2 \text{ NaI} + \text{Pb}(\text{NO}_3)_2 \rightarrow 2 \text{ NaNO}_3 + \text{PbI}_2$
 - $\text{NaI} + \text{CS}_2 \rightarrow \text{CSi} + \text{Na}$
 - $\text{H}_2(\text{g}) + \text{CO} + \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{CO}_3$
 - $\text{Li}_3\text{PO}_4 \rightarrow 3 \text{ Li} + \text{P} + 2 \text{ O}_2(\text{g})$
 - $\text{CS}_2 + 2 \text{ F}_2(\text{g}) \rightarrow \text{CF}_4 + 2 \text{ S}$

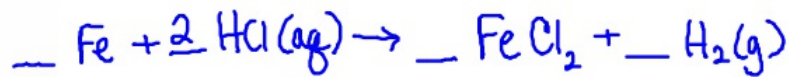
DR Predict the products for the following reactions, and write a balanced equation.



Dec
binary
chlorate
carbonate
hydroxide



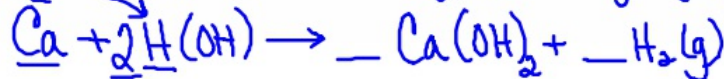
SR 11. iron + $\overset{+}{\text{H}} \overset{-}{\text{Cl}}$ hydrochloric acid \rightarrow iron(II) chloride + hydrogen gas



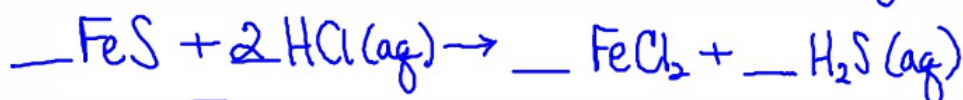
SR 14. magnesium bromide + chlorine gas \rightarrow magnesium chloride + bromine gas



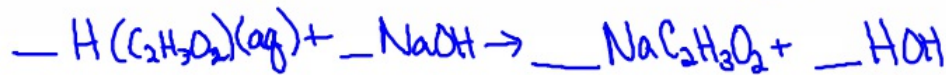
SR 15. calcium + water \rightarrow calcium hydroxide + hydrogen gas



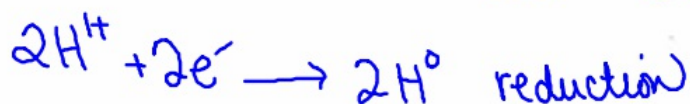
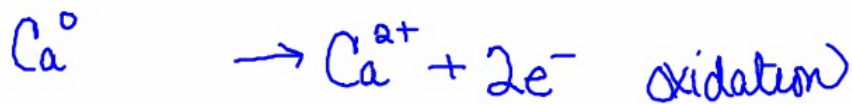
DR 17. iron(II) sulfide + hydrochloric acid \rightarrow iron(II) chloride + hydrosulfuric acid



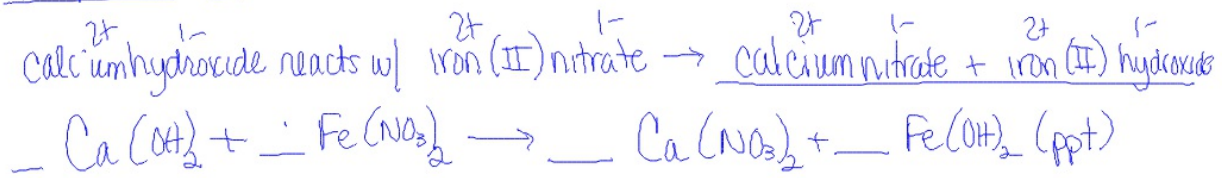
DR 19. acetic acid + sodium hydroxide \rightarrow sodium acetate + water



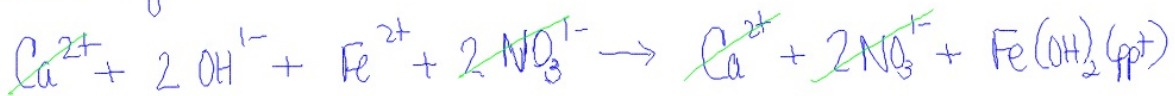
lose e^- oxidation gain e^- reduction



Precipitation (DR)



ionic equation



net ionic equation



Redox (DC or SR)

Zinc reacts with copper(II) chloride \rightarrow zinc chloride + copper

