

Name: _____ Block: ___ Date: _____

Nomenclature Review

Combine the following ions, (a.) record the formula and (b.) write the name of the compound formed.
(1/2 pt per box, 12 pts total)

		Br^{1-}	CO_3^{2-}	PO_4^{3-}
Zn^{2+}	a.			
	b.			
Mn^{4+}	a.			
	b.			
Fe^{3+}	a.			
	b.			
Li^{1+}	a.			
	b.			

Complete the following chart for the given components of the compound. (1/2 pt per box/ 12 pts total)

Example:			Formula	Name
Sodium	Chlorine	Na= +1 Cl= -1	NaCl	Sodium chloride

		Oxidation Number	Formula	Name
Ammonium	Nitrate	N = ___ H = ___ NH_4 N = ___ O = ___ NO_3		
Lead (II)	Phosphorous	Pb = ___ P = ___		
Silver	Chlorite	Ag = ___ Cl = ___ O = ___		
Calcium	Bicarbonate	Ca = ___ H = ___ C = ___ O = ___		
Chromium (III)	Sulfite	Cr = ___ S = ___ O = ___		
Magnesium	Oxygen	Mg = ___ O = ___		

Record the names of the following compounds. (3 pts each)

28. $\text{Al}_2(\text{CO}_3)_3$ _____

29. CO_2 _____

30. Fe_2O_3 _____

32. SF_2 _____

33. HCl (aq) _____

35. CoSO_4 _____

36. CuCl _____

38. $\text{H}(\text{C}_2\text{H}_3\text{O}_2)$ (aq) _____

Record the formula for the following compounds. (3 pts each)

39. Sulfur trioxide _____

40. Copper (II) oxide _____

41. Sodium sulfate _____

43. Mercury (II) nitrite _____

44. Dinitrogen pentaoxide _____

45. Calcium phosphate _____

48. Hydrobromic acid _____

50. Hydroiodic acid _____