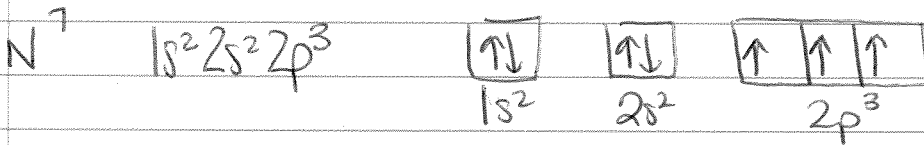
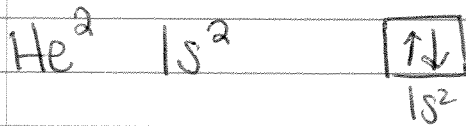
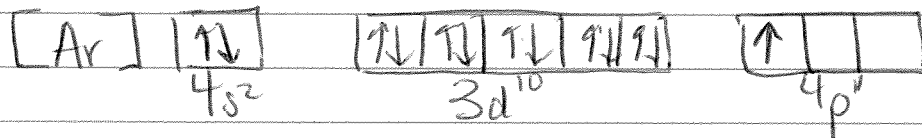
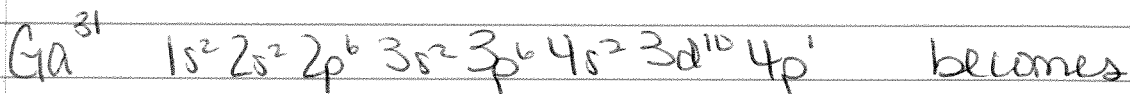
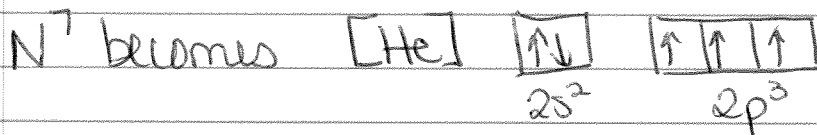


Examples:

Electron Configuration - placement of e^- w/out spin
Orbital Configuration - placement of e^- w/ spin



Shorthand Orbital Configuration - refer back to last completed noble gas and give orbital configuration for remaining e^-



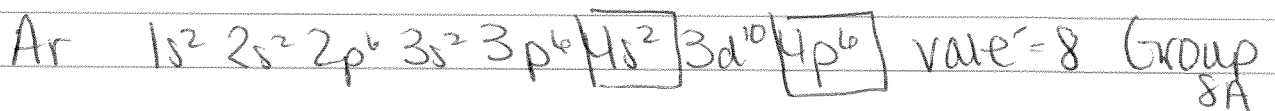
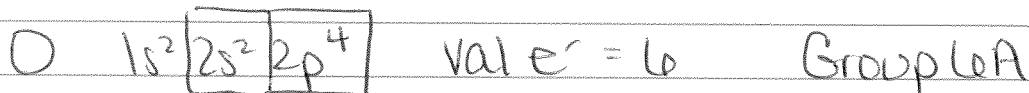
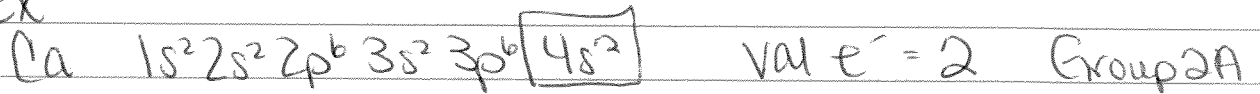
Practice: give all 3 versions for

Chlorine

Strontium

Valence Electrons: the outermost e^- (from the highest energy level) = represented by the highest s & p electrons.

EX



Practice: Complete electron conf. & determine val. e^-

① Be

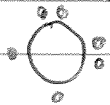
② Ni

③ Hg

Ni and Hg are both transition metals, what do you notice about their # of valence e^- ?

Lewis Dot Diagram: Represent val e^- as dots around the outside of the element symbol.

Ex. from above we know O has 6 val e^-



Practice: Complete the following Lewis Dot Diagrams:

① neon

② copper

③ Ba

④ uranium